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Chapter 11 Chemical Reactions 119 GUIDED PRACTICE PROBLEM 2 (page 324) 2. Sulfur burns in oxygen to form sulfur dioxide. Write a skeleton equation for this

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chemical reaction.

Include appropriate symbols from table 11.1. Analyze Step 1. Identify the relevant concepts. Write the formula for each reactant and each product. Include the common

SECTION 11.1
DESCRIBING
CHEMICAL
REACTIONS (pages
321-329)

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A representation of a chemical reaction with reactants on the left, products on the right, and an arrow separating the two.

Small whole number ratios that are placed in front of the formulas in an equation in order to balance it.

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Chemical
Reactions. In a
chemical reaction, the
reactants are written
on the left and the
products on the right.
The arrow that
separates them is
called yield.

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Chemical Reactions

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e The law of conservation of mass states that the mass of the reactants will always equal the mass of the products. In result the number of atoms of one element on the reactants side should be identical to the number of atoms of

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the same element on
the products side.

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Include the common

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PRACTICE PROBLEM 2

(page 324) 2. Sulfur

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burns in oxygen to form sulfur dioxide. Write a skeleton equation for this chemical reaction. Include appropriate symbols from table 11.1.

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Quantitative relationships exist with all chemical reactions that allow scientists to predict amounts of products formed, reactants consumed,

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and percent yield

based on theoretical

maximum. ... 15.2

CHAPTER 11:

STOICHIOMETRY. MOLE

TO MOLE RATIO. When

nitrogen and hydrogen

gas are heated under

the correct conditions,

ammonia gas (NH_3) is

formed.

CHAPTER 11:

STOICHIOMETRY

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Notes. Name: _____

11.1 - Describing

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Chemical Reactions. 1. What two substances reacted in the Hindenburg crash. 2. In general, what happens in a chemical reaction. 3. Look at Figure 11.1. What are some of the evidences of. chemical change (a.k.a. chemical reaction) 4. Give an example of a word equation for a chemical reaction (you

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Problems. When we
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carry out a reaction in

either an industrial
setting or a laboratory,
it is easier to work with
masses of substances
than with the numbers
of molecules or moles.

We will first present
the method used in
most other books for
converting from the
mass of any reactant
or product to the mass
of any other reactant

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or product using a
balanced chemical
equation outlined ...

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oxidation–reduction
reactions A chemical
reaction that exhibits a
change in the oxidation

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states of one or more elements in the reactants that has the general form oxidant + reductant \rightarrow reduced oxidant + oxidized reductant. The general forms of these five kinds of reactions are summarized in Table 11.6.1, along with examples of each.

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