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Chapter 11 Study Guide Stoichiometry

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CHAPTER 11 SECTIONS

1 Defining

Stoichiometry 2

Stoichiometric

Calculations 3 Limiting

Reactants 4 Percent

Yield LaunchLAB What

evidence can you

observe that a reaction

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has stopped? During a
chemical reaction,

reactants are

consumed as new

products form. In this

lab, you will look for

signs a chemical

reaction has stopped.

Steps in Stoichiometric

Calculations

CHAPTER 11

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outline of the chapter,
please use the book,
notes & homework to
study.) 11.1 Defining

Stoichiometry Vocab •

stoichiometry • mole

ratio Concepts Using

Balanced Equations •

Number of Atoms •

Number of Molecules •

Number of Moles •

Mass o Law of

Conservation of Mass •

Volume 11.2

Stoichiometric

Calculations Concepts

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Stoichiometry

368 Chapter 11 •

Stoichiometry Section

11.11.1 Objectives

Describe the types of relationships indicated by a balanced chemical equation. State the

mole ratios from a

balanced chemical

equation. Review

Vocabulary reactant:

the starting substance

in a chemical reaction

New Vocabulary

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stoichiometry mole
ratio Defining
Stoichiometry

Chapter 11:
Stoichiometry

In Section 11.3 , for example, you learned how to express the stoichiometry of the reaction for the ammonium dichromate volcano in terms of the atoms, ions, or molecules involved and the numbers of moles, grams, and formula

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units of each (recognizing, for instance, that 1 mol of ammonium dichromate produces 4 mol of water). This section describes how to use the stoichiometry of a reaction to answer questions like the following: How much oxygen is needed to ensure complete combustion of a ...

Chapter 11.4: **Stoichiometry -**

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Chemistry

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Stoichiometry Section
11.1 What is

stoichiometry? 1. true

2. true 3. false 4. true

5. true 6. 2, 2, 64.10 7.

3, 3, 96.00 8. 2, 2,

88.02 9. 4, 4, 72.08 10.

methanol and oxygen

gas 11. carbon dioxide

and water 12. 160.10 g

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Stoichiometry

13. 160.10 g 14. They are equal...

Stoichiometry
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Guide Answer Key

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Stoichiometry Section

11.1 What is

stoichiometry? 1. true

2. true 3. false 4. true

5. true 6. 2, 2, 64.10 7.

3, 3, 96.00 8. 2, 2,

88.02 9. 4, 4, 72.08 10.

methanol and oxygen

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gas 11. carbon dioxide
and water 12. 160.10 g
13. 160.10 g 14. They
are equal. 15. A mole
ratio is a ratio between
the numbers of moles

VIBRATIONS AND WAVES

CHAPTER 11:

STOICHIOMETRY. UNIT
4: Chemical Reactions,
The Mole,
Stoichiometry and
Thermodynamics. Part
B:Stoichiometry. Big
Picture Ideas; The

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Stoichiometry

identity of the reactants helps scientists to predict the products in a chemical reaction. Quantitative relationships exist with all chemical reactions that allow scientists to predict amounts of products formed, reactants consumed, and percent yield based on theoretical maximum.

**CHAPTER 11:
STOICHIOMETRY**

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Balanced Equations •

Number of Atoms •

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Chapter 5 - Mole

Concept and

Stoichiometry. chapter

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chapter 8 study guide

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$\tan A =$ Geometry - Ch.

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pursuing in CBSE Class

11 are advised

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**Chapter 11 Study
Guide Stoichiometry
Answer Key**

Chapter 11

Stoichiometry.

stoichiometry. mole
ratio. excess reactant.
limiting reactant. The
study of quantitative
relationships between
the amounts of.... In a
balanced equation, the
ratio between the
numbers of moles.... A
reactant that remains
after a chemical

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reaction stops.

Stoichiometry

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Chapter 11 The Mole
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Key Stoichiometry is
the tool for answering
these questions.

Stoichiometry The
study of quantitative
relationships between
the amounts of
reactants used and
amounts of products

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formed by a chemical reaction is called stoichiometry.

Stoichiometry is based on the law of conservation of mass.

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CHAPTER Section 11.1
continued In your
textbook, read about
mole ratios. Answer the
questions about the
following chemical
reaction. sodium +

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iron(III) oxide — ¥

sodium oxide + iron

$6\text{Na(s)} + \text{—+} + 2\text{Fe(s)}$

15. What is a mole ratio? 16. How IS a mole ratio written? CA S Q C CYA 17. Predict the number of mole ratios for this reaction. Class 18.

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CHAPTER 9 REVIEW
Stoichiometry SECTION
2 PROBLEMS Write the
answer on the line to

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the left. Show all your work in the space provided. 1. 4.5 mol
The following equation represents a laboratory preparation for oxygen gas: $2\text{KClO}_3(\text{s}) \rightarrow 2\text{KCl}(\text{s}) + 3\text{O}_2(\text{g})$ How many moles of O_2 form if 3.0 mol of KClO_3 are totally consumed? 2. 200 g
Given the ...

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